

by weight of substrate) at 40 atmospheres of pressure. The applicability and high efficiency of the recently developed noble metal synthetic high polymer catalysts, described in previous papers^{2,3,4,5} suggested that hydrogenation be attempted employing this type of catalyst. A palladium-polyvinyl alcohol (Pd-PVA) catalyst was prepared wherein the palladium was 0.5% by weight of substrate and cyanoacetic ethyl ester was successfully hydrogenated at atmospheric pressure.

Procedure

The Pd-PVA catalyst was composed of 30 mg. of Pd and 500 mg. of PVA in 100 cc. of 50% alcohol-water mixture which was 0.5 N in hydrochloric acid. The purpose of the high acidity was to minimize formation of secondary amines.⁶ Six and four-tenths g. of cyanoacetic ester was introduced into the vessel with the catalyst and the whole shaken for twenty hours at room temperature and ordinary pressure. At the end of this time 1822 cc. of a theoretical 2580 cc. of hydrogen had been adsorbed. The rate of adsorption was very slow at this point, so the isolation of β -alanine was started.

The colloidal catalyst was flocculated by addition of about 500 cc. of ethanol. After filtering, the liquid was concentrated. Hereafter, a separation into two layers was noticed, the lower layer containing some unreduced starting material, which is insoluble in water. The latter was drawn off and to assure complete removal of unreduced cyanoacetic ethyl ester, the remaining aqueous layer was extracted twice with ether. The aqueous solution was then taken to dryness under reduced pressure. The residual salt was twice extracted with 35-cc. portions of hot 95% ethanol, filtered while hot and allowed to cool. On cooling crystallization of ammonium chloride, which had been extracted along with the β -alanine ethyl ester hydrochloride, occurred. This was filtered off and discarded. The filtrate was again evaporated to dryness *in vacuo*. The residual salt was extracted with 35-cc. portions of absolute ethanol and filtered. The filtrate was concentrated *in vacuo* and then diluted with about 20 cc. of water and hydrolyzed under reflux. After concentrating under reduced pressure to remove the excess hydrochloric acid, the residue was dissolved in water and made up to a volume of 30 cc. A 2-cc. portion was withdrawn for determination of total halide.⁶ A suspension of silver oxide prepared from 10% more than the equivalent quantity of silver nitrate was added to the solution to bring about complete precipitation of halide.

After standing for some time, the precipitate was filtered, the filtrate concentrated and treated with hydrogen sulfide. The silver sulfide is centrifuged off, the solution treated with norite and filtered. The crystals of β -alanine may be obtained by concentrating the filtrate in a vacuum desiccator until crystallization begins. Upon recrystallization from water a product of m. p. 194–196° was obtained with the analysis calculated: N, 15.73; found: N, 15.56; yield, 15%.

No attempt was made to increase this yield as the primary purpose of the experiments was to demonstrate the ability of the colloidal noble metal synthetic high polymer catalysts to serve in the reduction of the nitrile group.

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(2) Louis D. Rampino and F. F. Nord, *THIS JOURNAL*, **68**, 2745, 3268 (1941).

(3) Rampino and Nord, *ibid.*, **65**, 429 (1943).

(4) Kevin E. Kavanagh, *ibid.*, **64**, 2721 (1942).

(5) Kevin E. Kavanagh and F. F. Nord, *ibid.*, **65**, 2121 (1943).

(6) "Organic Syntheses," Vol. 16, John Wiley and Son, Inc., New York, N. Y., 1936, p. 1.

sophical Society. Analyses were performed by Mr. Joseph Alicino of the Squibb Institute for Medical Research, New Brunswick, N. J.

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Chlorination of Fluorene with Sulfuryl Chloride

BY ANDREW STREITWIESER

Sulfuryl chloride is being used increasingly as a chlorinating agent because of the ease with which it reacts and the high yields obtained. In 1939, Kharasch and Brown¹ treated fluorene with this reagent and reported that the fluorene was chlorinated in the nucleus. However, they did not report the position taken by the entering chlorine atom, nor did they give any yields or procedures. The purpose of this work was to determine the position of the entering chlorine and to find out whether this method is a convenient one for preparing the chloro compound.

The fluorene was chlorinated in ether solution with a slight excess of sulfuryl chloride. The product melted at 95–96°.² This compares favorably with the 2-chlorofluorene prepared simultaneously by Chanussot³ and Courtot and Vignati,⁴ m. p. 96–97°. Authentic 2-chlorofluorene was prepared by the method of Chanussot³ from 2-aminofluorene by the Sandmeyer reaction. A mixed melting point determination showed the two to be identical. The following procedure is recommended as a quick, convenient method for preparing 2-chlorofluorene.

Experimental

2-Chlorofluorene from SO₂Cl₂.—Sixteen grams of fluorene and 120 cc. of anhydrous ether was placed in a distilling flask. Not all of the fluorene dissolved. The flask was stoppered with a rubber stopper carrying a separatory funnel containing 8 cc. (13.4 g.) sulfuryl chloride. The latter was added rapidly to the solution. When all was added, the ether was distilled off, and the residue recrystallized from alcohol. It can also be purified by steam distillation: yield, 85% of a white crystalline powder, m. p. 95–96°.

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(1) Kharasch and Brown, *THIS JOURNAL*, **61**, 2149 (1939).

(2) All melting points corrected.

(3) Chanussot, *Annales asocn. quim. Argentina*, **15**, 216 (1927).

(4) Courtot and Vignati, *Compt. rend.*, **184**, 1479 (1927).

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The Preparation of N-Alkyl Derivatives of *p*-Aminobenzoic Acid¹

BY A. R. SURREY AND H. F. HAMMER

It has been reported that procaine hydrochloride and other local anesthetics which are

(1) Presented before the Division of Medicinal Chemistry, A. C. S., Cleveland, Ohio, April, 1944.